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(54) **Zinc sulfophosphate glasses.**

(57) This invention relates to the preparation of a glass exhibiting a transition temperature normally below about 300 °, a working temperature below about 400 °, while, at the same time, exhibiting excellent resistance to attack by water. Specifically, the present invention discloses a glass consisting essentially in terms of mole percent on the oxide basis, of 15-35% P<sub>2</sub>O<sub>5</sub>, 1-25% SO<sub>3</sub>, 30-55% ZnO, 0-25% R<sub>2</sub>O, wherein R<sub>2</sub>O is selected from the group consisting of 0-25% Li<sub>2</sub>O, 0-25% Na<sub>2</sub>O, and 0-25% K<sub>2</sub>O, and up to a total of 15% of optional ingredients in the indicated proportions selected from the group consisting of 0-10% Al<sub>2</sub>O<sub>3</sub>, 0-10% MgO, 0-10% CaO, 0-10% SrO, 0-10% BaO, 0-10% MnO, 0-10% transition metal oxides and 0-15% Cl+F, as analyzed in weight percent.

**EP 0 648 713 A1**

FIELD OF THE INVENTION

The present invention relates to a zinc-sulfophosphate glass exhibiting an extremely low softening point of about less than 300 ° C and exhibiting excellent resistance to moisture attack.

BACKGROUND OF THE INVENTION

Considerable research has been conducted in the past to devise inorganic glasses exhibiting low transition temperatures ( $T_g$ ), thereby enabling melting and forming operations to be carried out at low temperatures with consequent savings in energy costs. As conventionally defined, the transition temperature of a glass is that temperature at which notable increase in thermal expansion coefficient is recorded, accompanied by a change in specific heat ( $C_p$ ). More recently, it has been recognized that glasses demonstrating low transition temperatures are potentially useful materials for a host of applications including low temperature sealing glasses and glass-organic polymer composites. A very recent development disclosed in U. S. Patent No. 5,043,369 (Bahn et al.) involves the production of glass-organic polymer alloys. Those alloys are prepared from a glass and a thermoplastic or thermosetting polymer having compatible working temperatures. Thus, the glass and the polymer are combined at the working temperature to form an intimate mixture; that is, the glass and polymer are in a sufficiently fluid state to be co-formed together to form a body displaying an essentially uniform, fine-grained microstructure in which, desirably, there is at least partial miscibility and/or a reaction between the glass and the polymer to promote adhesion and bonding therebetween. An article is shaped from the blend and then cooled to room temperature. Such articles exhibit chemical and physical properties comprising a complementary blend of those demonstrated by the particular glass and polymer. For example, the alloys frequently display a combination of high surface hardness, high stiffness, and high toughness. Alloys are distinguished from glass/organic polymer composites in that there is at least partial miscibility and/or a reaction between the glass and polymer which condition(s) is absent in glass/polymer composites.

Glasses having base compositions within the general zinc phosphate system have been found to be especially suitable for the glass component of glass-polymer alloys, U. S. Patent No. 4,940,677 and 5,071,795.

These zinc phosphate glasses demonstrate relatively good resistance to chemical attack when compared to other phosphate-based glasses. Nevertheless, the search has been continuous to discover new glass compositions manifesting low transition temperatures with even greater chemical durability. Not only would these lower temperature glass compositions result in lower energy costs attributed the glass formation, they would also result in lower costs attributed to formation of glass/polymer alloys and composites. In addition, lower temperature durable glasses would also increase the number of compatible polymers available which could be co-processed with the glass to form glass/polymer composites and thermally co-deformed with the glass to form glass/polymer alloys. These factors, lowered formation cost and increase in potential polymers choices, would, in turn, likely increase the number of potential commercial applications.

U.S. Patent No. 4,544,695 (Meyers) discloses a flame and/or smoke retardant polymeric composition which includes a phosphate-sulfate glass composition containing the following components: 4-18%  $K_2SO_4$ , 8-36%  $ZnSO_4$ , 4-18%  $Na_2SO_4$ , and 19-56%  $P_2O_5$ . In addition, the glass may include the following optional components: 0-25%  $ZnO$ , 0-4%  $B_2O_3$ , 0-25%  $Li_2O$ , 0-25%  $Na_2O$ , 0-12%  $BaO$  and 0-4%  $TiO_2$ . The inventive glasses, when expressed in the composition form used in the Meyers patent, fall outside the claimed composition range disclosed therein, and particularly in all three glass examples, the total amount of  $ZnO$  (combined from sources of  $As_2SO_4$  and  $ZnO$ ) is far below that required in the present inventive compositions; the highest  $ZnO$  cited is 18.5%. Even if the zinc oxide level claimed therein is raised enough to fall within the present inventive glasses, the required ranges of the Meyers patent will produce either a glass too high in sulfate to exhibit good durability, or too low in alkali to result in a glass with a low enough glass transition temperature. Since the glasses of the Meyers patent have either extremely poor durability or too high a transition temperature they are unsuited for use in glass/polymer alloys and composites.

SUMMARY OF THE INVENTION

The above described research has led to the invention of a glass exhibiting a transition temperature normally below about 300 °, a working temperature below about 400 °, while, at the same time, exhibiting excellent resistance to attack by water.

More specifically, the present invention discloses a glass consisting essentially in terms of mole percent on the oxide basis, of 15-35%  $P_2O_5$ , 1-25%  $SO_3$ , 30-55%  $ZnO$ , 0-25%  $R_2O$ , wherein  $R_2O$  is selected from the group consisting of 0-25%  $Li_2O$ , 0-25%  $Na_2O$ , and 0-25%  $K_2O$ . Additionally, up to a total of 15% of optional ingredients in the indicated proportions may be selected from the group consisting of 0-10%  $Al_2O_3$ , 0-10%  $MgO$ , 0-10%  $CaO$ , 0-10%  $SrO$ , 0-10%  $BaO$ , 0-10%  $MnO$ , 0-10% transition metal oxides and 0-15%  $Cl + F$ , as analyzed in weight percent.

Whereas complete conversion of the above inventive composition ranges expressed in terms of mole percent to ranges expressed in terms of weight percent is not mathematically possible, the following are the inventive compositions as expressed in terms of approximate weight percent: 25-50%  $P_2O_5$ , 1-20%  $SO_3$ , 26-45%  $ZnO$ , 0-20%  $R_2O$ , wherein  $R_2O$  is selected from the group consisting of 0-8%  $Li_2O$ , 0-20%  $Na_2O$ , and 0-20%  $K_2O$ . The group from which the optional ingredients may be selected from, as expressed in terms of approximate weight percent, is as follows: 0-10%  $Al_2O_3$ , 0-5%  $MgO$ , 0-5%  $CaO$ , 0-10%  $SrO$ , 0-15%  $BaO$ , 0-10%  $MnO$ , 0-15% transition metal oxides and 0-15%  $Cl + F$ , as analyzed.

## DESCRIPTION OF THE PREFERRED EMBODIMENTS

Table I records a number of glass compositions, expressed in terms of mole percent on the oxide basis, operable in the present invention. Table IA reports the same compositions, but wherein the values of the individual components have been converted to weight percent. In addition to reporting the relative amounts of the batch constituents, Table I reports the transition temperature ( $T_g$ ) in terms of  $^{\circ}C$ , as measured by employing standard differential scanning calorimetry techniques, the dissolution rate expressed in terms of percentage weight loss (% wt loss) of a tab of approximate dimensions of about 1" x 1½" x ⅜" (2½cm x 4cm x 1cm) and the working temperature (pull temp.). In addition, the appearance of the glass after formation is reported in Table I (Glass appear.); the quality and appearance of the glass formed ranged from clear (cl.) to slightly hazy (hazy). The actual batch ingredients for the glasses can comprise any materials, either the oxides or other compounds, which, upon being melted together, will be converted into the desired oxides in the proper proportions. For example  $ZnSO_4 \cdot 7H_2O$  is conveniently employed as the source of  $SO_3$  as well as a partial source of  $ZnO$ .

The batch materials were compounded, automatically tumble-mixed in order to secure a homogeneous melt, and placed into silica crucibles. The glass batch was then melted and maintained at temperatures ranging from 800 to 1000° for times of about three hours. Very little volatilization of  $P_2O_5$ ,  $SO_3$  or other species was noted below about 850° C. Table I reports that the analyzed values of  $SO_3$  (A- $SO_3$ ) were typically about equal to that calculated from the batch (B- $SO_3$ ); values reported in weight percent. The melt was then poured into a steel mold to produce a rectangular slab having dimensions of about 4" x 8" x ⅜" (10cm x 20 x 1cm) and that slab subsequently annealed overnight at temperature of about 250 to 275° C.

Rectangular tab-shaped pieces weighing approximately 40 grams were cut from the slab, heated to temperatures within the range of 350 to 450° C, and glass cane was hand drawn to obtain a close approximation of the working temperature. This value, for glass Examples 1-14, is reported in Table I as the "Pull temp."

Samples were cut from the glass slab for testing the moisture resistance/durability thereof. The test involved weighing each sample carefully and then immersing the sample into a bath of boiling water. After a residence time of six hours, the sample was removed from the bath, dried in the ambient environment, and thereafter reweighed to determine any loss of weight by the sample. This loss of weight, i.e., the dissolution rate, for each glass sample is reported in Table I (Dissol. rate) and is calculated as a percentage of the original untested/unimmersed weight.

TABLE I

	1	2	3	4	5	6	7	8	9	10	11	12
P <sub>2</sub> O <sub>5</sub>	26.0	22.7	19.5	24.3	26.4	24.8	23.1	26.4	25.0	25.0	30.0	23.0
SO <sub>3</sub>	10.0	15.0	20.0	15.7	10.0	12.5	15.0	15.0	15.0	15.0	12.5	12.5
Al <sub>2</sub> O <sub>3</sub>	1.6	1.4	1.2	1.4	1.6	1.5	1.4	1.6	---	---	2.0	1.5
Li <sub>2</sub> O	6.4	5.6	4.8	5.8	5.6	5.3	4.9	5.6	---	---	5.0	5.5
Na <sub>2</sub> O	9.7	9.5	9.3	9.0	8.0	8.0	8.0	8.0	---	---	8.0	8.0
K <sub>2</sub> O	6.4	6.7	6.9	6.0	5.7	5.8	6.0	5.7	---	---	4.0	5.5
ZnO	38.1	36.9	35.7	35.9	41.6	40.7	39.9	40.7	45.0	40.0	38.5	42.5
CaO	1.3	1.5	1.6	1.3	0.6	0.7	1.0	0.6	---	---	---	0.75
SrO	0.6	0.8	1.1	0.5	0.6	0.7	1.0	0.6	---	---	---	0.75
ZnCl <sub>2</sub>	---	---	---	---	---	---	---	---	15.0	20.0	---	---
Melt. temp. (°C)	800	800	800	800	800	800	800	800	--	1000	850	825
Glass Appear.	cl.	hazy	cl.	hazy	cl.	cl.	cl.	cl./ hazy	cl.	cl.	cl.	cl.
T <sub>g</sub> (°C)	--	--	--	288	286	276	283	--	--	--	--	--
Pull (°C)	385	--	375	380	400	385	375	390	--	--	400	405
Dissolution Rate (% wt. loss)	0.02	0	0.15	0.03	0	0	0.12	0.0	--	--	0.0	0.02
A - SO <sub>3</sub>	8.7	13.4	18.0	13.4	--	--	--	--	--	--	--	--
B - SO <sub>3</sub>	8.6	13.2	17.9	12.9	--	--	--	--	--	--	--	--

TABLE IA

	<u>1</u>	<u>2</u>	<u>3</u>	<u>4</u>	<u>5</u>	<u>6</u>	<u>7</u>	<u>8</u>	<u>9</u>	<u>10</u>	<u>11</u>	<u>12</u>
P <sub>2</sub> O <sub>5</sub>	39.8	33.3	30.9	37.5	39.9	37.8	35.6	38.3	34.0	33.1	44.3	35.6
SO <sub>3</sub>	8.6	13.2	17.9	13.7	8.5	10.7	13.0	12.2	11.5	11.2	10.4	10.9
Al <sub>2</sub> O <sub>3</sub>	1.7	1.5	1.3	1.6	1.7	1.6	1.5	1.6	-	-	2.1	1.6
Li <sub>2</sub> O	2.0	1.9	1.6	1.9	1.8	1.7	1.6	1.7	-	-	1.6	1.7
Na <sub>2</sub> O	6.5	6.5	6.5	6.1	5.3	5.4	5.4	5.1	-	-	5.2	5.4
K <sub>2</sub> O	6.5	6.9	7.3	6.1	5.7	5.9	6.2	5.5	-	-	4.0	5.7
ZnO	33.4	32.9	32.4	31.8	36.1	35.6	35.3	34.6	35.0	30.3	32.5	37.7
CaO	0.8	0.9	1.0	0.8	0.3	0.4	0.4	0.3	-	-	-	0.4
SrO	0.6	0.9	1.2	0.6	0.6	0.8	0.9	0.6	-	-	-	0.9
ZnCl <sub>2</sub>	--	--	--	--	--	--	--	-	19.5	25.4	-	-

Whereas the above description reflects laboratory melting and forming practice only, it will be appreciated that the recited compositions are capable of being melted in large scale melting units and shaped into desired configurations utilizing forming techniques conventional in the glassmaking art. As is the case with standard glassmaking practice, it is only necessary to ensure that the batch materials are mixed together thoroughly and then melted at temperatures which will ensure a homogenous melt without

excessive volatilization of sulfide oxides, chloride and fluoride, and that the melt is thereafter cooled and shaped into a glass body of a desired geometry which is customarily annealed.

These inventive glasses possessing transition temperatures normally below about 300 °C and working temperatures of 350 ° and above, exhibit measures of durability/resistance to moisture ranging from 0.0 to 0.15 percent weight loss when immersed in boiling water for six (6) hours. These weight loss values are exceptionally low for glasses with transition/working temperatures in the range which these inventive glasses exhibit; they are comparable to those measures of durability exhibited by glasses possessing transition/working temperatures as much as 100 °C greater, e.g., those phosphate glasses disclosed in U.S. Pat. No. 4,940,677 (Beall et al.)

Based upon an overall balance of physical and chemical properties, Example 6 is the most preferred embodiment of the inventive glasses.

Table II reports the composition of three typical metaphosphate glasses, exhibiting transition temperatures comparable to those exhibited by the inventive glasses. It is clear from the durability data reported therein (only a 1 hour immersion in boiling water), ranging from a 5% weight loss to a complete dissolving of the glass, that the inventive glasses exhibit a much greater durability than these metaphosphate samples possessing comparable transition temperatures.

TABLE II

	13	14	15
P <sub>2</sub> O <sub>5</sub>	46	46	46
Al <sub>2</sub> O <sub>3</sub>	4	4	4
Li <sub>2</sub> O	25	--	25
Na <sub>2</sub> O	25	25	--
K <sub>2</sub> O	--	25	25
Dissolution Rate (% wt. loss)	5.0	dissolved (100%)	25
T <sub>g</sub> ( ° C)	290	260	280

Table III reports the composition of Examples 1-12 in the form as is used in the earlier-described Meyers reference. In converting the instant examples to this form, it was assumed that first source of the SO<sub>3</sub> was K<sub>2</sub>SO<sub>4</sub>; if any SO<sub>3</sub> was still needed it was then assumed that it was supplied through the addition of Na<sub>2</sub>SO<sub>4</sub>; and finally if an amount of SO<sub>3</sub> was still required it was assumed that its source was ZnSO<sub>4</sub>. This order of preference for designating the sulfate in terms of K<sub>2</sub>SO<sub>4</sub>, Na<sub>2</sub>SO<sub>4</sub>, and ZnSO<sub>4</sub>, is clear from Examples 1 and 2 of the Meyers patent. In Example 1, all of the alkali, K<sub>2</sub>O and Na<sub>2</sub>O, is listed in terms of sulfates. The zinc by comparison, is listed as both ZnSO<sub>4</sub> and ZnO. In other words, zinc is designated as sulfate up to the total sulfate level desired by the patentee, and the rest of the zinc is listed as the oxide. This could just as easily been accomplished by using all ZnSO<sub>4</sub> and listing part of the Na or K as oxides, but the patentee evidently preferred to list the alkali ahead of zinc as sulfates. Similarly, from Example 2, all the potassium is listed as K<sub>2</sub>SO<sub>4</sub>, while both soda and zinc are split between oxides and sulfates. While any arrangement can obviously be used, the patentee prefers K, Na, and Zn as the order of preference in describing his compositions. We have used the same format (order of preference) in Table III.

It is clear from the compositions listed in this table that the inventive glasses are compositionally very distinct from those disclosed in the Meyers et al. reference.

TABLE III

	1	2	3	4	5	6	7	8	9	10	11	12
$K_2SO_4$	7.1	7.9	8.7	7.1	6.3	6.6	6.7	6.4	-	-	4.6	6.3
$Zn_2SO_4$	-	-	4.7	0.8	-	-	1.1	1.3	17.6	17.6	0.6	-
$Na_2SO_4$	4.0	9.8	11.7	10.7	4.8	7.6	8.9	9.0	-	-	9.2	8.0
$P_2O_5$	28.9	26.7	24.0	28.9	29.3	28.3	25.7	29.6	29.4	29.4	34.3	26.3
ZnO	42.3	43.4	40.1	41.8	46.2	46.5	48.3	44.2	35.3	29.4	43.4	48.6
$Li_2O$	7.1	6.6	6.0	6.9	6.2	6.1	5.5	6.3	-	-	5.7	6.3
$Na_2O$	6.8	1.4	-	-	4.1	1.5	-	-	-	-	-	-
CaO	1.4	1.8	2.0	1.5	0.7	0.8	1.1	0.7	-	-	-	0.9
SrO	0.7	0.9	1.4	0.6	0.7	0.8	1.1	0.7	-	-	-	0.9
$ZnCl_2$	-	-	-	-	-	-	-	-	17.6	23.5	-	-
$Al_2O_3$	1.8	1.6	1.5	1.7	1.8	1.7	1.6	1.8	-	-	2.3	1.7

Table IV reports a number of glass compositions expressed in terms of mole percent on the oxide basis which are outside the scope of the instant invention, but within the composition range disclosed in the earlier-described Meyers reference. In addition to the composition breakdown and the dissolution rate of the glass, the table reports the melting temperature (Melt. temp.) and the glass appearance, before and after the immersion in the boiling water. The glass, before immersion, ranges in appearance and quality from a clear stable glass (cl.) to a translucent hygroscopic glass (transl./hygro). Table V reports the same

compositions though reported in the composition form used in the Meyers patent. These comparison glass examples, 16 to 22 were, except for changing the respective components as shown in Table II, made in the same way as that earlier disclosed for inventive glass Examples 1-12. Note, however, that Example 20 (reported as unmelted) would not melt, thus no glass was formed; this composition would not melt to form a glass even when the melting instructions detailed in the Meyers reference were adhered to.

It is quickly evident that, with the exception of Example 21, these glasses exhibit poor durability measures; either high dissolution rates or hygroscopic behavior. In fact, this durability is poor enough to effectively render these glasses useless for any application envisioned for the inventive glasses. It is important to note that these weight loss percentages are values for various immersion times; Ex. 16 and 17 - 2½ hrs., Ex. 19 - 2 hrs, and Ex. 18 and 20-23 - 6 hrs. As reported in Table II, the examples exhibited weight losses as a result of boiling water immersion ranging from a 0.3% weight loss in 6 hours, to those which either dissolved (dissol.) or were converted into a unconsolidated white residue (unconsol. white res.) or possessed surface spalling (sur. spall.)

Examples 16-18 are specific examples disclosed in the Meyers patent as possessing low hydroscopic behavior; however, the durability values measured show that none of these compositions exhibit or approach the durability as exhibited by the inventive composition. Examples 19 through 23 were various composition attempts in order to identify a glass composition within this reference which would exhibit durability on the order of the inventive sulfophosphate composition's durability. For example, comparison examples 21 and 22 possesses the maximum amount of Zn (highest ZnSO<sub>4</sub> and ZnO allowable) and the lowest amount of P<sub>2</sub>O<sub>5</sub> allowable within the claimed composition range; neither exhibit durability numbers as low as desired. Example 21, while approaching the low dissolution rates of the current invention, is nevertheless far too high in transition temperature (>350 °C) and working temperature (>450 °C) to be useful in the current invention. Simply put, none of the comparison examples, and likely no composition falling within the Meyers reference composition range possesses the requisite high ZnO and sufficiently low SO<sub>3</sub> in order to ensure a glass which possesses a T<sub>g</sub> of about 300 °C with a corresponding excellent resistance to moisture.



TABLE IV

	<u>16</u>	<u>17</u>	<u>18</u>	<u>19</u>	<u>20</u>	<u>21</u>	<u>22</u>	<u>23</u>
P <sub>2</sub> O <sub>5</sub>	26.4	22.7	45.6	56	19	21.5	12.2	26.2
SO <sub>3</sub>	28.0	38.8	17.4	16	16	30.5	35.9	18.0
K <sub>2</sub> O	7.0	5.9	4.4	4	4	2.8	6.4	3.3
Na <sub>2</sub> O	7.0	17.1	4.4	7	29	2.8	6.5	14.8
Li <sub>2</sub> O	13.2	---	6.5	---	15	--	--	5.7
ZnO	18.4	15.7	8.7	33	33	42.4	39.1	32.0
B <sub>2</sub> O <sub>3</sub>	---	---	3.2	---	---	--	--	--
BaO	---	---	9.7	---	---	--	--	--
Melt temp. (°C)	600-700	600-700	600-750	600-750	unmelted	875	800	800
Glass appear.	cl.	cl.	transl./hygro.	cl./hygro.	--	cl.	cl.	cl.
T <sub>g</sub> (°C)	233	249	264	243	--	>350	--	--
Pull (°C)					--	>450	375	375
Dissolution rate (% wt. loss)	40	80	0.2	51	51	0.2	100	1.0
Appearance after boil	unconsol. white res.	unconsol. white res.	hazy glass	dissol.	unmelted	--	unconsol. white res.	sur. spall.

TABLE V

<u>16</u>	<u>17</u>	<u>18</u>	<u>19</u>	<u>20</u>	<u>21</u>	<u>22</u>	<u>23</u>
9.7	9.7	5.3	4.0	4.0	4.0	10.0	4.0
19.5	25.6	10.6	8.0	8.0	36.0	36.0	14.0
9.7	28.0	5.3	4.0	4.0	4.0	10.0	4.0
36.7	36.7	55.3	56.0	19.0	31.0	19.0	32.0
6.1	--	--	25.0	25.0	25.0	25.0	25.0
18.3	--	7.9	--	15.0	--	--	7.0
--	--	--	3.0	25.0	--	--	14.0
--	--	11.8	--	--	--	--	--
--	--	3.9	--	--	--	--	--
K <sub>2</sub> SO <sub>4</sub>							
Zn <sub>2</sub> SO <sub>4</sub>							
Na <sub>2</sub> SO <sub>4</sub>							
P <sub>2</sub> O <sub>5</sub>							
ZnO							
Li <sub>2</sub> O							
Na <sub>2</sub> O							
BaO							
B <sub>2</sub> O <sub>3</sub>							

**Claims**

1. A glass exhibiting a transition temperature below about 300° , a working temperature below about 400° and exhibiting excellent resistance to attack by water and mild aqueous alkaline solutions, the

glass consisting essentially, expressed in terms of mole percent on the oxide basis, of 15-35%  $P_2O_5$ , 1-25%  $SO_3$ , 30-55%  $ZnO$ , 0-25%  $R_2O$ , wherein  $R_2O$  is selected from the group consisting of 0-25%  $Li_2O$ , 0-25%  $Na_2O$ , and 0-25%  $K_2O$ , and up to a total of 15% of optional ingredients in the indicated proportions selected from the group consisting of 0-10%  $Al_2O_3$ , 0-10%  $MgO$ , 0-10%  $CaO$ , 0-10%  $SrO$ , 0-10%  $BaO$ , 0-10%  $MnO$ , 0-10% transition metal oxides and 0-15%  $Cl + F$ , as analyzed in weight percent.

2. The glass as according to claim 1 wherein the amount of  $SO_3$  ranges from 7-20%.

3. The glass according to claim 1 wherein the amount of  $SO_3$  does not exceed 15%.

4. The glass according to claim 1 consisting essentially of 21-33%  $P_2O_5$ , 9-17%  $SO_3$ , 35-51%  $ZnO$ , and 10-20%  $R_2O$ , wherein  $R_2O$  is selected from the group consisting of 0-10%  $Li_2O$ , 0-15%  $Na_2O$ , and 0-10%  $K_2O$ .

5. The glass according to claim 1 consisting essentially of 24.8%  $P_2O_5$ , 12.5%  $SO_3$ , 40.7%  $ZnO$ , 5.3%  $Li_2O$ , 8.0%  $Na_2O$ , 5.8%  $K_2O$ , 1.5%  $Al_2O_3$ , 0.7%  $CaO$ , and 0.7%  $SrO$ .



European Patent  
Office

## EUROPEAN SEARCH REPORT

Application Number  
EP 94 11 5569

DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION
D,A	US-A-4 544 695 (MYERS) * column 3, line 27 - column 4, line 25; examples * ---	1-5	C03C3/16 C03C3/17 C03C14/00 C08K3/40
A	DATABASE WPI Section Ch, Week 8837, Derwent Publications Ltd., London, GB; Class L01, AN 88-262869 & SU-A-1 377 251 (URALS KIROV POLY) 29 February 1988 * abstract *	1-5	
A	DATABASE WPI Section Ch, Week 8806, Derwent Publications Ltd., London, GB; Class H07, AN 88-041156 & SU-A-1 318 561 (URALS KIROV POLY) 23 June 1987 * abstract *	1-5	
A D,A	EP-A-0 365 236 (CORNING INCORPORATED) & US-A-5 043 369 -----		TECHNICAL FIELDS SEARCHED (Int. Cl.-6)  C03C
The present search report has been drawn up for all claims			
Place of search THE HAGUE		Date of completion of the search 17 January 1995	Examiner Van Bommel, L
<b>CATEGORY OF CITED DOCUMENTS</b> X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure P : intermediate document  T : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application L : document cited for other reasons ----- & : member of the same patent family, corresponding document			